



Class 11 Chemistry | Some Basic Concepts Of Chemistry

MCQ WORKSHEET • 15 questions • Choose one correct option • Suggested time: 25 minutes

1. How many moles are present in 13.2 g CO₂? (M = 44 g mol⁻¹)

A. 0.60 B. 0.15 C. 0.30 D. 3.0

2. How many oxygen atoms are present in 0.25 mol CO₂?

A. 0.50 N_a B. 1.00 N_a C. 0.25 N_a D. 0.75 N_a

3. What is the mass of 0.20 mol NaOH? (M = 40 g mol⁻¹)

A. 2 g B. 20 g C. 8 g D. 80 g

4. A compound has empirical formula CH₂ and molar mass 56. Its molecular formula is:

A. C₄H₁₀ B. C₄H₈ C. C₂H₄ D. C₃H₆

5. What is the molarity of 0.050 mol solute in 250 mL solution?

A. 2.0 M B. 0.0125 M C. 0.20 M D. 0.50 M

6. 100 mL of 0.60 M solution is diluted to 300 mL. New molarity is:

A. 0.020 M B. 0.60 M C. 1.80 M D. 0.20 M

7. For N₂ + 3H₂ → 2NH₃, 2 mol N₂ and 3 mol H₂ give at most:

A. 3 mol NH₃ B. 1 mol NH₃
C. 2 mol NH₃ D. 4 mol NH₃

8. Mass percentage of oxygen in H₂O is approximately:

A. 50.0% B. 66.7% C. 88.9% D. 11.1%

9. How many significant figures are in 0.00450?

A. 4 B. 5 C. 3 D. 2

10. Correctly reported sum of 12.11 g and 0.3 g is:

A. 12.410 g B. 12.4 g C. 12.41 g D. 12 g

11. The mole fraction of A in 2 mol A mixed with 3 mol B is:

A. 0.67 B. 1.50 C. 0.60 D. 0.40

12. A 10 g solute is dissolved in 90 g water. Mass percentage is:

A. 9% B. 11.1% C. 90% D. 10%

13. Which sample contains the most molecules?

A. 44 g CO₂ B. 28 g N₂ C. 8 g H₂ D. 16 g O₂

14. A sample has 24 g C and 4 g H. Its empirical formula is:

A. CH₄ B. C₂H C. C₂H₃ D. CH₂

15. The same mass of C combines with 16 g and 32 g O in two oxides. This illustrates:

A. Conservation of volume
B. Reciprocal proportions only
C. Constant gas pressure
D. Multiple proportions

ANSWER KEY — check after completing all 15 questions

1.C 2.A 3.C 4.B 5.C 6.D 7.C 8.C 9.C 10.B 11.D 12.D 13.C 14.D 15.D

Original exam-style practice • Reference notes: www.physics.guru/cbse-practice-papers/ • Not an official paper



Class 11 Chemistry | Structure Of Atom

MCQ WORKSHEET • 15 questions • Choose one correct option • Suggested time: 25 minutes

1. Maximum electrons in the $n = 3$ shell are:

A. 32 B. 8 C. 9 D. 18

2. Which quantum-number pair is forbidden?

A. $n=3, l=2$ B. $n=1, l=0$
C. $n=2, l=2$ D. $n=2, l=1$

3. How many orbitals belong to a d subshell?

A. 10 B. 5 C. 7 D. 3

4. The number of radial nodes in a 3p orbital is:

A. 2 B. 1 C. 3 D. 0

5. A photon of frequency ν has energy:

A. $h\nu$ B. $h\nu$ C. ν/h D. h/ν

6. If light wavelength doubles, photon energy becomes:

A. Half B. Unchanged C. Double D. Four times

7. An electron's momentum doubles. Its de Broglie wavelength becomes:

A. Unchanged B. Four times C. Half D. Double

8. Which transition in hydrogen emits the highest-energy photon?

A. $3 \rightarrow 2$ B. $4 \rightarrow 3$ C. $5 \rightarrow 4$ D. $3 \rightarrow 1$

9. Hydrogen's Balmer series terminates at:

A. $n=3$ B. $n=4$ C. $n=1$ D. $n=2$

10. Which orbital is spherical?

A. 2s B. $2p_x$ C. $3d_{yz}$ D. $3d_{xy}$

11. The ground-state configuration of Cr ($Z=24$) is:

A. $[\text{Ar}]3d^5 4s^1$ B. $[\text{Ar}]3d^6$
C. $[\text{Ar}]3d^4 4s^2$ D. $[\text{Ar}]3d^3 4s^3$

12. In ground-state nitrogen, the number of unpaired electrons is:

A. 5 B. 1 C. 2 D. 3

13. Two electrons in one orbital must have:

A. Different l B. Different n
C. Opposite spins D. Same spin

14. In a many-electron atom, which orbital fills before 3d by the Aufbau rule?

A. 5p B. 4d C. 4s D. 4p

15. An atom has $Z=17$ and mass number 37. Its neutron count is:

A. 17 B. 54 C. 37 D. 20

ANSWER KEY — check after completing all 15 questions

1.D 2.C 3.B 4.B 5.B 6.A 7.C 8.D 9.D 10.A 11.A 12.D 13.C 14.C 15.D

Original exam-style practice • Reference notes: www.physics.guru/cbse-practice-papers/ • Not an official paper

Plot No. 15, Second Floor, Jaisingh Chowk, Sector 85, Gurgaon - 122004

8800 8800 28 | www.goodmarksclasses.com



Class 11 Chemistry | Classification Of Elements And Periodicity In Properties

MCQ WORKSHEET • 15 questions • Choose one correct option • Suggested time: 25 minutes

1. Which has the largest atomic radius?

A. Mg B. Si C. Al D. Na

2. Which species has the smallest radius?

A. F^- B. Na^+ C. Mg^{2+} D. Al^{3+}

3. Which element has the highest electronegativity?

A. F B. Cl C. O D. N

4. First ionisation enthalpy of Mg exceeds that of Al because Mg has:

A. More protons
B. A filled 3s subshell
C. A smaller nuclear charge than Na
D. A larger 3p population

5. First ionisation enthalpy of N exceeds that of O mainly due to:

A. Filled $2p^6$ in N
B. Larger N atom
C. Greater shielding in N
D. Stable half-filled $2p^3$

6. Which has the most negative electron-gain enthalpy among these?

A. I B. Br C. F D. Cl

7. An element ending in $3s^23p^4$ belongs to:

A. Group 14, period 3 B. Group 6, period 2
C. Group 16, period 4 D. Group 16, period 3

8. Across a period, effective nuclear charge generally:

A. Remains zero B. Decreases
C. Increases D. Becomes negative

9. Down group 1, first ionisation enthalpy generally:

A. Becomes zero B. Increases
C. Remains constant D. Decreases

10. Which oxide is most strongly basic?

A. Al_2O_3 B. SO_3 C. Na_2O D. SiO_2

11. Which oxide is amphoteric?

A. Na_2O B. Al_2O_3 C. CO_2 D. SO_2

12. An element with $Z=20$ belongs to which block?

A. f B. d C. p D. s

13. Modern periodic law relates properties primarily to:

A. Atomic number B. Mass number
C. Neutron number D. Atomic mass only

14. Which atom forms the largest anion among these?

A. Cl B. Br C. I D. F

15. Which pair contains isoelectronic species?

A. Ca^{2+} and Mg^{2+} B. Na and Ne
C. O^{2-} and Mg^{2+} D. Cl^- and F^-

ANSWER KEY — check after completing all 15 questions

1.D 2.D 3.A 4.B 5.D 6.D 7.D 8.C 9.D 10.C 11.B 12.D 13.A 14.C 15.C

Original exam-style practice • Reference notes: www.physics.guru/cbse-practice-papers/ • Not an official paper

Plot No. 15, Second Floor, Jaisingh Chowk, Sector 85, Gurgaon - 122004

8800 8800 28 | www.goodmarksclasses.com



Class 11 Chemistry | Chemical Bonding And Molecular Structure

MCQ WORKSHEET • 15 questions • Choose one correct option • Suggested time: 25 minutes

1. Shape of NH_3 is:

- A. Trigonal pyramidal B. Linear
C. Tetrahedral D. Trigonal planar

2. Which molecule has zero net dipole moment?

- A. H_2O B. NH_3 C. CO_2 D. SO_2

3. Central-atom hybridisation in BF_3 is:

- A. sp B. sp^3d C. sp^2 D. sp^3

4. Bond angle in gaseous BeCl_2 is:

- A. 90° B. 120° C. 109.5° D. 180°

5. Which bond is formed by sideways p-orbital overlap?

- A. π bond B. Hydrogen bond
C. Ionic bond D. σ bond

6. Number of σ and π bonds in ethene is:

- A. 4 and 2 B. 5 and 1 C. 3 and 3 D. 6 and 0

7. O_2 is paramagnetic because it contains:

- A. A lone pair on one atom only
B. Only paired electrons
C. Two unpaired electrons
D. No antibonding electrons

8. MO bond order of N_2 is:

- A. 3 B. 2 C. 1 D. 0

9. Compared with O_2 , O_2^+ has:

- A. Same bond order B. Higher bond order
C. Lower bond order D. Zero bond order

10. Formal charge on N in NH_4^+ is:

- A. 0 B. +1 C. -1 D. +2

11. Strongest repulsion in VSEPR is between:

- A. Two nuclei only
B. A bond pair and nucleus
C. Two lone pairs
D. Two bond pairs

12. Which substance shows extensive intermolecular hydrogen bonding?

- A. CH_4 B. PH_3 C. H_2S D. H_2O

13. Shape of SF_6 is:

- A. Trigonal bipyramidal B. Tetrahedral
C. Square planar D. Octahedral

14. Shape of XeF_4 is:

- A. Square planar B. Tetrahedral
C. See-saw D. Trigonal pyramidal

15. Which has greatest covalent character by Fajans' rule?

- A. LiF B. KF C. NaF D. LiI

ANSWER KEY — check after completing all 15 questions

1.A 2.C 3.C 4.D 5.A 6.B 7.C 8.A 9.B 10.B 11.C 12.D 13.D 14.A 15.D

Original exam-style practice • Reference notes: www.physics.guru/cbse-practice-papers/ • Not an official paper



Class 11 Chemistry | Thermodynamics

MCQ WORKSHEET • 15 questions • Choose one correct option • Suggested time: 25 minutes

1. A system absorbs 120 J heat and does 50 J work. ΔU is:

- A. 170 J B. 70 J C. -170 J D. -70 J

2. For expansion against constant external pressure, work on the system is:

- A. $+P_{ext} \Delta V$ B. Zero in every case
C. $-P_{ext} \Delta V$ D. $P_{ext}/\Delta V$

3. Which is a state function?

- A. Heat B. Path length C. Work D. Enthalpy

4. At constant volume with only PV work, heat equals:

- A. ΔU B. $-\Delta U$ C. $T\Delta S$ always D. ΔH always

5. For an ideal-gas reaction, $\Delta H - \Delta U$ equals:

- A. $-\Delta n_g RT$ B. $RT/\Delta n_g$ C. $\Delta n_g RT$ D. Zero always

6. For $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$, Δn_g is:

- A. -2 B. -1 C. +1 D. +2

7. Standard formation enthalpy of $O_2(g)$ in its reference state is:

- A. 0 B. 1 kJ mol^{-1}
C. Always positive D. -1 kJ mol^{-1}

8. If $A \rightarrow B$ has $\Delta H = -40 \text{ kJ}$, then $B \rightarrow A$ has ΔH :

- A. 0 B. $+40 \text{ kJ}$ C. -40 kJ D. $+80 \text{ kJ}$

9. $A \rightarrow B$: $+25 \text{ kJ}$; $B \rightarrow C$: -70 kJ . $A \rightarrow C$ is:

- A. $+45 \text{ kJ}$ B. -95 kJ C. -45 kJ D. $+95 \text{ kJ}$

10. Which process usually increases entropy?

- A. Condensing steam B. Crystallising a salt
C. Freezing water D. Melting ice

11. At constant T, P a spontaneous change has:

- A. $\Delta G > 0$ B. $\Delta G < 0$
C. $\Delta G = 0$ always D. $\Delta S_{\text{system}} < 0$ always

12. For $\Delta H = -20 \text{ kJ}$ and $\Delta S = -50 \text{ J K}^{-1}$, ΔG at 300 K is:

- A. -35 kJ B. $+35 \text{ kJ}$ C. -5 kJ D. $+5 \text{ kJ}$

13. A reaction with $\Delta H > 0$ and $\Delta S > 0$ is favoured at:

- A. No temperature
B. Low temperature only
C. High temperature
D. All temperatures

14. For a reversible isothermal process, entropy change is:

- A. $q_{rev} T$ B. $-q_{rev}/T$ always C. T/q_{rev} D. q_{rev}/T

15. If a reaction has $\Delta G^\circ = 0$, its equilibrium constant is:

- A. 10 B. Infinity C. 0 D. 1

ANSWER KEY — check after completing all 15 questions

1.B 2.C 3.D 4.A 5.C 6.A 7.A 8.B 9.C 10.D 11.B 12.C 13.C 14.D 15.D

Original exam-style practice • Reference notes: www.physics.guru/cbse-practice-papers/ • Not an official paper



Class 11 Chemistry | Equilibrium

MCQ WORKSHEET • 15 questions • Choose one correct option • Suggested time: 25 minutes

1. For $\text{H}_2 + \text{I}_2 \rightleftharpoons 2\text{HI}$, the equilibrium expression K_c is:

- A. $[\text{H}_2][\text{I}_2]/[\text{HI}]^2$ B. $[\text{HI}]^2/[\text{H}_2]$
C. $[\text{HI}]^2/([\text{H}_2][\text{I}_2])$ D. $[\text{HI}]/([\text{H}_2][\text{I}_2])$

2. At equilibrium, forward and reverse reaction rates are:

- A. Equal and nonzero B. Both zero always
C. Reverse greater D. Forward greater

3. For $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$, higher pressure favours:

- A. Reactants B. No shift
C. Only N_2 removal D. NH_3 formation

4. For an exothermic forward reaction, raising temperature:

- A. Decreases K B. Makes K zero
C. Increases K D. Leaves K unchanged

5. A catalyst added to an equilibrium mixture:

- A. Raises K
B. Lowers K
C. Leaves K unchanged
D. Stops reverse reaction

6. If K_c for $\text{A} \rightleftharpoons \text{B}$ is 4, K_c for $\text{B} \rightleftharpoons \text{A}$ is:

- A. 0.25 B. 4 C. -4 D. 16

7. For a reaction with $\Delta n_g = 0$, K_p is:

- A. K_c B. $K_c RT$ C. K_c^2 D. K_c/RT

8. At 25°C , pH of 0.001 M HCl is approximately:

- A. 7 B. 1 C. 11 D. 3

9. At 25°C , pH of 0.010 M NaOH is approximately:

- A. 12 B. 2 C. 7 D. 10

10. The conjugate base of H_2PO_4^- is:

- A. H_3O^+ B. H_3PO_4 C. HPO_4^{2-} D. PO_4^{3-}

11. Which pair forms an acidic buffer?

- A. $\text{HNO}_3/\text{KNO}_3$
B. $\text{CH}_3\text{COOH}/\text{CH}_3\text{COONa}$
C. NaOH/NaCl
D. HCl/NaCl

12. For weak acid HA with $C=0.10$ M and $K_a=10^{-5}$, $[\text{H}^+]$ is approximately:

- A. 10^{-5} M B. 10^{-6} M C. 10^{-2} M D. 10^{-3} M

13. If solubility of AgCl is s mol L^{-1} in pure water, K_{sp} is:

- A. s^2 B. $4s^3$ C. s D. $2s^2$

14. Adding NaCl to saturated AgCl solution generally:

- A. Increases AgCl solubility
B. Reduces AgCl solubility
C. Leaves solubility unchanged
D. Makes K_{sp} zero

15. Which species acts as a Lewis acid?

- A. Cl^- B. NH_3 C. BF_3 D. OH^-

ANSWER KEY — check after completing all 15 questions

1.C 2.A 3.D 4.A 5.C 6.A 7.A 8.D 9.A 10.C 11.B 12.D 13.A 14.B 15.C



Class 11 Chemistry | Redox Reactions

MCQ WORKSHEET • 15 questions • Choose one correct option • Suggested time: 25 minutes

1. Oxidation number of Mn in KMnO_4 is:

A. +4 B. +2 C. +6 D. +7

2. Oxidation number of sulfur in SO_4^{2-} is:

A. +6 B. +4 C. -2 D. +8

3. Oxidation number of oxygen in H_2O_2 is:

A. 0 B. +1 C. -1 D. -2

4. In $\text{Zn} + \text{Cu}^{2+} \rightarrow \text{Zn}^{2+} + \text{Cu}$, the reducing agent is:

A. Cu^{2+} B. Zn^{2+} C. Zn D. Cu

5. Oxidation is associated with:

- A. Electron loss
- B. Electron gain
- C. Decrease in oxidation number
- D. Proton gain only

6. Which change is a reduction?

- A. $\text{Cl}^- \rightarrow \text{Cl}_2$
- B. $\text{Zn} \rightarrow \text{Zn}^{2+}$
- C. $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+}$
- D. $\text{Fe}^{3+} \rightarrow \text{Fe}^{2+}$

7. Electrons accepted by one MnO_4^- to form Mn^{2+} in acid are:

A. 5 B. 2 C. 7 D. 3

8. Electrons lost when one Cr^{2+} becomes Cr^{3+} are:

A. 2 B. 1 C. 5 D. 3

9. Which reaction is disproportionation?

- A. $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
- B. $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$
- C. $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
- D. $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

10. Oxidation number of H in NaH is:

A. +2 B. 0 C. +1 D. -1

11. In $\text{Cl}_2 + 2\text{OH}^- \rightarrow \text{Cl}^- + \text{ClO}^- + \text{H}_2\text{O}$, chlorine changes from 0 to:

- A. -2 and +2
- B. -1 and +1
- C. -1 and +5
- D. +1 and +3

12. In acid, $\text{Cr}_2\text{O}_7^{2-}$ needs how many electrons to form 2Cr^{3+} ?

A. 12 B. 2 C. 3 D. 6

13. Which is not a redox reaction?

- A. $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$
- B. $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$
- C. $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$
- D. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$

14. For $\text{MnO}_4^- + \text{Fe}^{2+}$ in acid, the mole ratio $\text{MnO}_4^- : \text{Fe}^{2+}$ is:

A. 1:5 B. 2:1 C. 1:2 D. 5:1

15. The average oxidation number of S in $\text{S}_2\text{O}_3^{2-}$ is:

A. +2 B. +6 C. +4 D. -2

ANSWER KEY — check after completing all 15 questions

1.D 2.A 3.C 4.C 5.A 6.D 7.A 8.B 9.C 10.D 11.B 12.D 13.B 14.A 15.A

Original exam-style practice • Reference notes: www.physics.guru/cbse-practice-papers/ • Not an official paper

Plot No. 15, Second Floor, Jaisingh Chowk, Sector 85, Gurgaon - 122004

8800 8800 28 | www.goodmarksclasses.com



Class 11 Chemistry | Organic Chemistry: Some Basic Principles And Techniques

MCQ WORKSHEET • 15 questions • Choose one correct option • Suggested time: 25 minutes

1. IUPAC name of $\text{CH}_3\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_3$ is:

- A. 3-Methylbutane B. 2-Methylbutane
C. Pentane D. 2-Ethylpropane

2. Heterolytic cleavage of a covalent bond produces:

- A. Only molecules B. Two free radicals
C. Only neutral atoms D. Ions

3. Homolytic cleavage produces:

- A. Only carbocations B. Only carbanions
C. Ionic crystals D. Free radicals

4. Which is a nucleophile?

- A. BF_3 B. H^+ C. NO_2^+ D. OH^-

5. Which is an electrophile?

- A. NH_3 B. NO_2^+ C. OH^- D. CN^-

6. Most stable simple alkyl carbocation is:

- A. CH_3^+ B. CH_3CH_2^+ C. $(\text{CH}_3)_3\text{C}^+$ D. $(\text{CH}_3)_2\text{CH}^+$

7. A carbon atom in a $\text{C}\equiv\text{C}$ bond is typically:

- A. Unhybridised B. sp^3 hybridised
C. sp^2 hybridised D. sp hybridised

8. Ethanol and dimethyl ether show:

- A. Geometrical isomerism
B. Functional isomerism
C. Chain isomerism only
D. Optical isomerism

9. Which group generally shows a -I effect?

- A. $-\text{CH}_3$ B. $-\text{C}(\text{CH}_3)_3$ C. $-\text{NO}_2$ D. $-\text{C}_2\text{H}_5$

10. Resonance structures differ in:

- A. Electron arrangement
B. Molecular formula
C. Atomic positions
D. Total charge

11. A sublimable solid can be separated from nonvolatile impurities by:

- A. Decantation alone B. Filtration alone
C. Centrifugation alone D. Sublimation

12. Miscible liquids with close boiling points are separated by:

- A. Sublimation B. Fractional distillation
C. Crystallisation only D. Simple filtration

13. Chromatography separates substances by differences in:

- A. Nuclear charge alone
B. Combustion temperature
C. Distribution between phases
D. Total atom count only

14. In sodium fusion, nitrogen in an organic compound is converted to:

- A. NaNO_3 B. N_2O C. NaCN D. NH_4Cl

15. How many chain isomers does C_5H_{12} have?

- A. 2 B. 3 C. 5 D. 4

ANSWER KEY — check after completing all 15 questions

1.B 2.D 3.D 4.D 5.B 6.C 7.D 8.B 9.C 10.A 11.D 12.B 13.C 14.C 15.B



Class 11 Chemistry | Hydrocarbons

MCQ WORKSHEET • 15 questions • Choose one correct option • Suggested time: 25 minutes

1. Chlorination of methane in sunlight proceeds mainly by:

- A. Electrophilic addition
- B. Nucleophilic addition
- C. Free-radical substitution
- D. Condensation

2. Ethene reacts with Br₂ in an inert solvent to form:

- A. Bromoethene
- B. Ethane
- C. 1,2-Dibromoethane
- D. Bromoethane

3. Major product of propene with HBr, without peroxide, is:

- A. 2-Bromopropane
- B. 1-Bromopropane
- C. 1,2-Dibromopropane
- D. Propane

4. Propene with HBr in peroxide gives mainly:

- A. 2-Bromopropane
- B. Propane
- C. 1,1-Dibromopropane
- D. 1-Bromopropane

5. Which hydrocarbon gives a precipitate with ammoniacal AgNO₃?

- A. Propane
- B. Propyne
- C. Propene
- D. But-2-yne

6. Hydrogenation of ethyne using Lindlar catalyst gives:

- A. Ethanol
- B. Ethanal
- C. Ethane
- D. Ethene

7. Reductive ozonolysis of propene gives:

- A. Methanal and ethanal
- B. Ethanoic acid and methane
- C. Only methanal
- D. Only propanone

8. Benzene undergoes nitration using:

- A. NaNO₂ and water
- B. Dilute HNO₃ only
- C. NH₃ and O₂
- D. Conc. HNO₃ and conc. H₂SO₄

9. The electrophile in benzene nitration is:

- A. NO₂⁺
- B. NO₃⁻
- C. NO₂⁻
- D. NH₄⁺

10. Which group is meta-directing in electrophilic aromatic substitution?

- A. -NH₂
- B. -NO₂
- C. -OH
- D. -CH₃

11. Complete combustion of 1 mol propane consumes:

- A. 3 mol O₂
- B. 4 mol O₂
- C. 5 mol O₂
- D. 7 mol O₂

12. Which compound can show cis-trans isomerism?

- A. But-1-ene
- B. But-2-ene
- C. 2-Methylpropene
- D. Ethene

13. Most stable conformation of ethane is:

- A. Staggered
- B. Planar trans
- C. Planar cis
- D. Eclipsed

14. Sodium propanoate heated with soda lime yields:

- A. Ethane
- B. Methane
- C. Ethene
- D. Propane

15. Benzene's unusual stability is explained by:

- A. Delocalised π electrons
- B. sp³ carbon atoms
- C. Six ionic bonds
- D. Three isolated double bonds

ANSWER KEY — check after completing all 15 questions

1.C 2.C 3.A 4.D 5.B 6.D 7.A 8.D 9.A 10.B 11.C 12.B 13.A 14.A 15.A